

Determining Atomic Mass Worksheet 1

The following elements are shown with all stable, non-trace isotopes. Use the mass numbers values, not atomic mass unit values, along with the relative abundances provided, to determine the atomic mass of each element. Compare your calculated number to that in the periodic table to determine if your estimated answer is realistic. **You must set up the equation and show all work to receive credit.**

1. Oxygen

Oxygen-16	99.76%
Oxygen-17	0.039%
Oxygen-18	0.201%

Atomic Mass - O

2. Lithium

Lithium-6	7.5%
Lithium-7	92.5%

Atomic Mass - Li

3. Nitrogen

Nitrogen-14	99.634%
Nitrogen-15	0.366%

Atomic Mass - N

4. Titanium

Titanium-46	8.0%
Titanium-47	7.3%
Titanium-48	73.8%
Titanium-49	5.5%
Titanium-50	5.4%

Atomic Mass - Ti

5. Germanium

Germanium-70	21.23%
Germanium-72	27.66%
Germanium-73	7.73%
Germanium-74	35.94%
Germanium-76	7.44%

Atomic Mass - Ge